



INDIAN SCHOOL AL WADI AL KABIR



DEPARTMENT OF SCIENCE

Sample Question Paper 2 (TERM – I)

2021-22

Class X

Science (086)

**Time: 90 Minutes**

**General Instructions:**

1. The Question Paper contains three sections.
2. Section A has 24 questions. Attempt any 20 questions.
3. Section B has 24 questions. Attempt any 20 questions.
4. Section C has 12 questions. Attempt any 10 questions.
5. All questions carry equal marks.
6. There is no negative marking.

**SECTION -  
A**

Section – A consists of 24 questions. Attempt any 20 questions from this section.

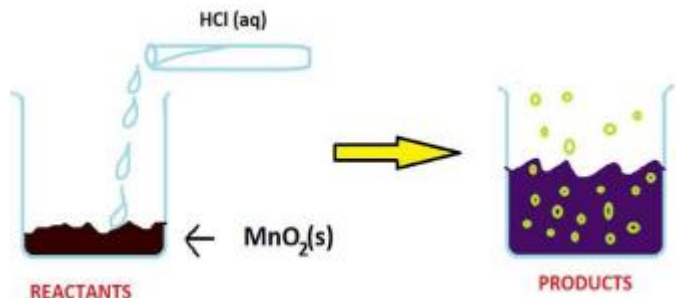
The first attempted 20 questions would be evaluated.

|    |  |
|----|--|
| 1. | What change in colour is observed when white silver chloride is left exposed to sunlight?<br>(a) White to brown<br>(b) White to grey<br>(c) White to green<br>(d) No change in colour.   |
| 2. | The chemical reaction between copper (II)oxide and hydrogen is given below.<br>$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$<br>The reducing and oxidising agents involved in this redox reaction are:<br>(a) Hydrogen and copper oxide respectively<br>(b) Hydrogen and copper respectively<br>(c) Copper and hydrogen respectively<br>(d) Hydrogen and water respectively |
| 3. | If a sample of water containing detergents is provided to you, which of the following methods will you adopt to neutralise it?<br>(a) Treating the water with baking soda<br>(b) Treating the water with vinegar<br>(c) Treating the water with caustic soda   |

(d) Treating the water with washing soda

4.

The reaction of  $\text{MnO}_2$  with  $\text{HCl}$  is depicted in the following diagram. It was observed that a gas with bleaching abilities was released.



In the above discussed reaction, what is the nature of  $\text{MnO}_2$ ?

- (a) Acidic oxide
- (b) Basic oxide
- (c) Neutral oxide
- (d) Amphoteric oxide

5.

$\text{ZnO} + (\text{X}) \rightarrow \text{ZnSO}_4 + \text{H}_2\text{O}$ . Here (X) is-

- (a)  $\text{ZnSO}_4$
- (b)  $\text{HCl}$
- (c)  $\text{H}_2\text{SO}_4$
- (d)  $\text{HNO}_3$

6.

Plaster of Paris hardens by-

- (a) giving off  $\text{CO}_2$ .
- (b) changing into  $\text{CaCO}_3$ .
- (c) combining with water
- (d) giving out water

7.

When zinc granules are dipped in blue coloured copper sulphate solution, the colour of solution obtained is

- (a) blue
- (b) yellow
- (c) colourless
- (d) brown

8.

A student tests a sample of drinking water and reports its pH value as 5 at room temperature. Which one of the following might have been added in water?

- (a) Ammonium chloride
- (b) Sodium chloride
- (c) Sodium bicarbonate
- (d) Baking soda

|     |   |
|-----|---|
| 9.  | <p>Which of the following statements is not correct?</p> <p>(i) All metal carbonates react with acid to give a salt, water and carbon dioxide<br/> (ii) All metal oxides react with water to give salt and acid<br/> (iii) Some metals react with acids to give salt and hydrogen<br/> (iv) Some non-metallic oxides react with water to form an acid</p> <p>(a) (i)            (b) (ii)            (c) (iii)            (d) (iv)</p> |
| 10. | <p>The colour of coating developed on a zinc rod on dipping it in an aqueous copper sulphate solution will be</p> <p>(a) Blue            (b) Brown            (c) White            (d) Green</p>  |
| 11. | <p>Select the correct statement.</p> <p>(a) Heterotrophs do not synthesis their own food<br/> (b) Heterotrophs utilise solar energy for photosynthesis<br/> (c) Heterotrophs synthesise their own food<br/> (d) Heterotrophs are capable of converting carbon dioxide and water into carbohydrates.</p>   |
| 12. | <p>What is observed when air is blown from mouth into a test tube containing lime water?</p> <p>(a) Lime water turns milky due to the CO<sub>2</sub> exhaled<br/> (b) Lime water becomes colourless due to the CO<sub>2</sub> exhaled<br/> (c) Lime water turns milky due to water vapour in blown air<br/> (d) Lime water does not turn milky due to the CO<sub>2</sub> exhaled</p>  |
| 13. | <p>The dialyser works as kidney except does not perform.....</p> <p>(a) osmoregulation<br/> (b) tubular secretion<br/> (c) selective reabsorption<br/> (d) ultrafiltration</p>  |
| 14. | <p>The blood leaving the tissues becomes richer in</p> <p>(a) Carbon dioxide<br/> (b) Water<br/> (c) Haemoglobin<br/> (d) Oxygen</p>  |
| 15. | <p>The role of nasal cavity in human respiratory system</p> <p>(i) Filtration of inhaled air.<br/> (ii) Removal of germs and dust.<br/> (iii) Moistening of the inhaled air.</p> <p>(a)(i) &amp;(ii)<br/> (b)(ii) &amp; (iii)<br/> (c)(i), (ii) &amp; (iii)</p>   |

(d) None of these

16.

Which of the following statement(s) is (are) true about respiration?

- (i) During inhalation, ribs move inward and diaphragm is raised
  - (ii) In the alveoli, exchange of gases takes place i.e., oxygen from alveolar air diffuses into blood and carbon dioxide from the blood into the alveolar air
  - (iii) Haemoglobin has a greater affinity for carbon dioxide than oxygen
  - (iv) Alveoli increase surface area for exchange of gases
- (a) (i) and (iv)  
(b) (ii) and (iii)  
(c) (i) and (iii)  
(d) (ii) and (iv)

17.

In an experiment to trace the path of a ray of light passing through a rectangular glass slab, four students tabulated their observations as given below. Which student performed the experiment most correctly?

| Student    | A          | B            | C          | D            |
|------------|------------|--------------|------------|--------------|
| $\angle i$ | $30^\circ$ | $30^\circ$   | $30^\circ$ | $30^\circ$   |
| $\angle r$ | $18^\circ$ | $20^\circ$   | $19^\circ$ | $21.5^\circ$ |
| $\angle e$ | $32^\circ$ | $32.5^\circ$ | $30^\circ$ | $29^\circ$   |

- (a) A    (b) B    (c) C    (d) D

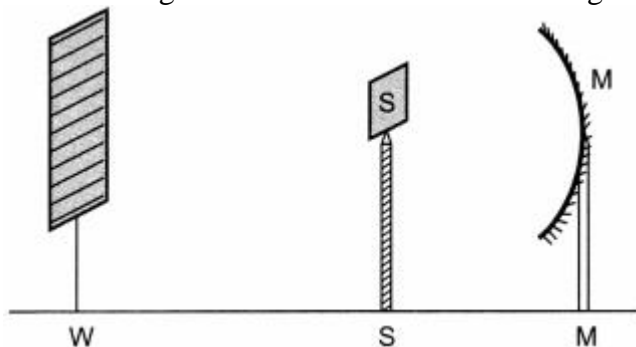
18.

What happens when white light is passed from air to glass prism

- (a) Bends away from normal
- (b) Bends towards normal
- (c) Passes un deviated
- (d) Reflects back

19.

A student obtains a sharp image of the distant window (W) of the school laboratory on the screen (S) using the given concave mirror (M) to determine its focal length. Which of the following distances should he measure to get the focal length of the mirror?

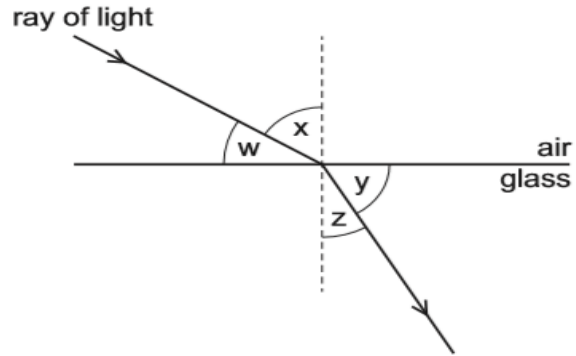


- (a) MW  
(b) MS

- (c) SW
- (d) MW- WS

20.

The diagram shows a ray of light passing from air into glass.

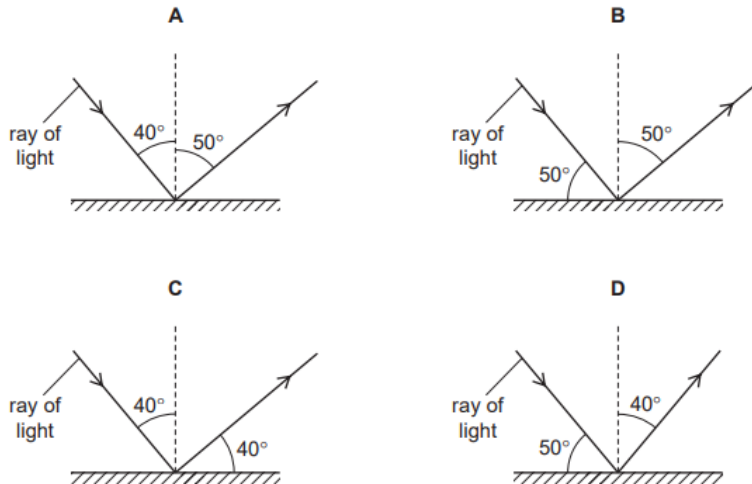


Which labelled angles are the angle of incidence and the angle of refraction?

|          | angle of incidence | angle of refraction |
|----------|--------------------|---------------------|
| <b>A</b> | w                  | y                   |
| <b>B</b> | w                  | z                   |
| <b>C</b> | x                  | y                   |
| <b>D</b> | x                  | z                   |

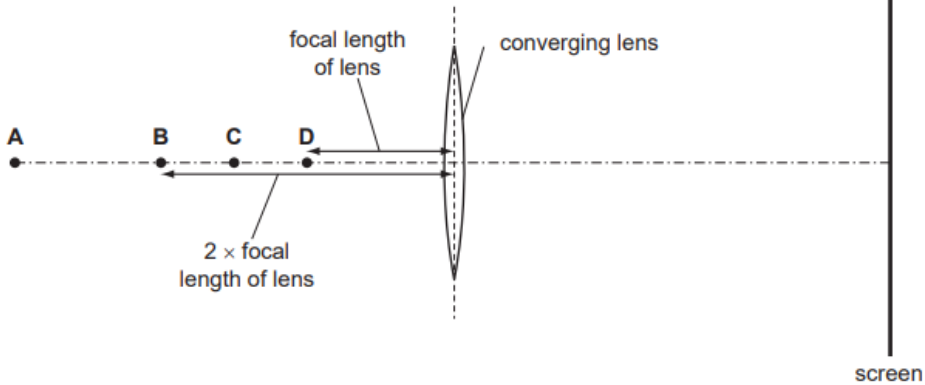
21.

Which diagram correctly shows a ray of light reflected by a plane mirror?

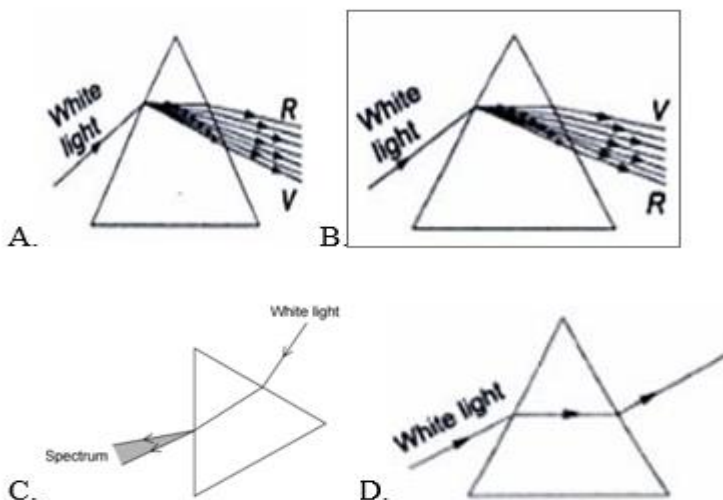


22.

A converging lens in a projector is used to make an enlarged image of a small piece of film on a screen. At which labelled point could the piece of film be placed so that the lens produces this image?



23. Which of the following figures correctly represents the passage of white light through prism?



24. If a man's face is 25 cm in front of concave shaving mirror producing erect image 1.5 times the size of face, focal length of the mirror would be \_\_\_\_\_  
 (a) -25 cm      (b) -75cm      (c) -50cm      (d) 25cm

**SECTION -**

**B**

Section - B consists of 24 questions (Sl. No.25 to 48). Attempt any 20 questions from this section. The first attempted 20 questions would be evaluated.

25. An iron nail was kept immersed in aluminium sulphate solution, after about an hour it was observed that:  
 (a) The colourless solution changes to light green  
 (b) the solution becomes warm  
 (c) grey metal is deposited on the iron nail  
 (d) the solution remains colourless and no deposition is seen on the iron nail.

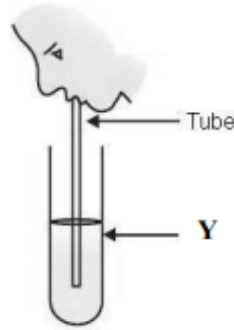
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| 26.  | $\text{Fe}_2\text{O}_3 + \text{Al} \rightarrow \text{Al}_2\text{O}_3 + \text{Fe}$ . The reaction is an example of:<br>(a) combination reaction<br>(b) decomposition reaction<br>(c) displacement reaction<br>(d) double displacement reaction   |
| 27.  | The reaction of water and quick lime is an example of:<br>(a) combination reaction<br>(b) exothermic reaction<br>(c) both (a) and (b)<br>(d) None of these.   |
| 28.  | A student strongly heats hydrated ferrous sulphate salt in a dry test tube. He would observe a:<br>(a) yellow residue<br>(b) brown residue<br>(c) light green residue<br>(d) white residue  |
| 29.  | A teacher gave two test tubes – one containing water and the other containing sodium hydroxide solution to two students. Then he asked them to identify the test tube containing sodium hydroxide solution. Which one of the following can be used for correctly identifying the test tube containing the solution of sodium hydroxide?<br>(a) Blue litmus<br>(b) Red litmus<br>(c) Sodium carbonate solution<br>(d) Dilute HCl |
| 30.  | Sodium carbonate is a basic salt because it is a salt of:<br>(a) Strong acid and strong base<br>(b) Weak acid and weak base<br>(c) Strong acid and weak base<br>(d) Weak acid and strong base   |
| <p><b>Question No. 31 to 35 consist of two</b> statements – Assertion (A) and Reason (R). Answer these questions selecting the appropriate option given below:</p> <p>A. Both A and R are true and R is the correct explanation of A<br/>           B. Both A and R are true and R is not the correct explanation of A<br/>           C. A is true but R is false<br/>           D. A is False but R is true</p> |   |
| 31.  | Assertion: Zinc oxide is an amphoteric oxide.<br>Reason: It reacts with both acid and base to form salt and water.  |

|     |  |
|-----|--|
| 32. | Assertion: Bread tastes sweet on mastication.<br>Reason: Salivary amylase converts starch into sugar.  |
| 33. | Assertion: mammals and birds have four chambered heart.<br>Reason: mammals and birds are warm blooded.   |
| 34. | Assertion(A): Virtual images are always erect.<br>Reason (R) : Virtual images are formed by converging lenses only   |
| 35. | Reddish brown deposits observed on iron nails when these are kept in aqueous solution of copper sulphate is that of:<br>(a) $\text{Cu}_2\text{O}$<br>(b) $\text{CuO}$<br>(c) $\text{Cu}$<br>(d) $\text{CuS}$   |
| 36. | The atomic numbers of four elements A, B, C and D are 6, 8, 10 and 12 respectively. The two elements which can react to form ionic bonds (or ionic compound) are:<br>(a) A and D (b) B and C (c) A and C (d) B and D   |
| 37. | Given below are the steps to be followed for performing starch test on a green leaf<br>(i) Boil the leaf in alcohol<br>(ii) Boil the leaf in water<br>(iii) Dip the leaf in iodine solution<br>(iv) Wash the leaf in water<br>The correct sequences of steps are<br>a) (i), (iv), (ii), (iii)<br>b) (ii), (iv), (iii), (i)<br>c) (ii), (i), (iii), (iv)<br>d) (iv), (i), (ii), (iii)   |
| 38. | Select the most appropriate sequence for aerobic respiration<br>a) $\text{Glucose} \xrightarrow{\text{Mitochondria}} \text{Pyruvate} \xrightarrow{\text{Cytoplasm}} \text{CO}_2 + \text{H}_2\text{O} + \text{Energy}$<br>b) $\text{Glucose} \xrightarrow{\text{Cytoplasm}} \text{Pyruvate} \xrightarrow{\text{Mitochondria}} \text{CO}_2 + \text{H}_2\text{O} + \text{Energy}$<br>c) $\text{Glucose} \xrightarrow{\text{Cytoplasm}} \text{Pyruvate} + \text{Energy} \xrightarrow{\text{Mitochondria}} \text{CO}_2 + \text{H}_2\text{O}$<br>d) $\text{Glucose} \xrightarrow{\text{Cytoplasm}} \text{Pyruvate} + \text{Energy} \xrightarrow{\text{Mitochondria}} \text{CO}_2 + \text{H}_2\text{O} + \text{Energy}$ |



39.

Observe the given diagram and answer the following question



To test the release of CO<sub>2</sub> gas during respiration, the chemical used in test tube (Y) is

- a) Lime
- b) Lime water
- c) Calcium carbonate
- d) Marble

40.

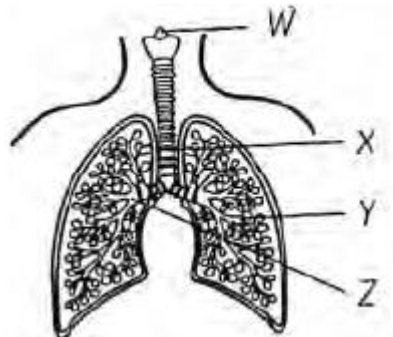
Match the words of Column (A) with that of Column (B)

| COLUMN (A) |           | COLUMN (B) |                       |
|------------|-----------|------------|-----------------------|
| a)         | Phloem    | (i)        | Excretion             |
| b)         | Nephron   | (ii)       | Translocation of food |
| c)         | Veins     | (iii)      | Clotting of blood     |
| d)         | Platelets | (iv)       | Deoxygenated blood    |

- a) a) (ii), b) (i), c) (iv) d) (iii)
- b) a) (iii), b) (i), c) (iv) d) (ii)
- c) a) (ii), b) (i), c) (iii) d) (iv)
- d) a) (ii), b) (iii), c) (iv) d) (i)

41.

The diagram shows part of the human gas exchange system.

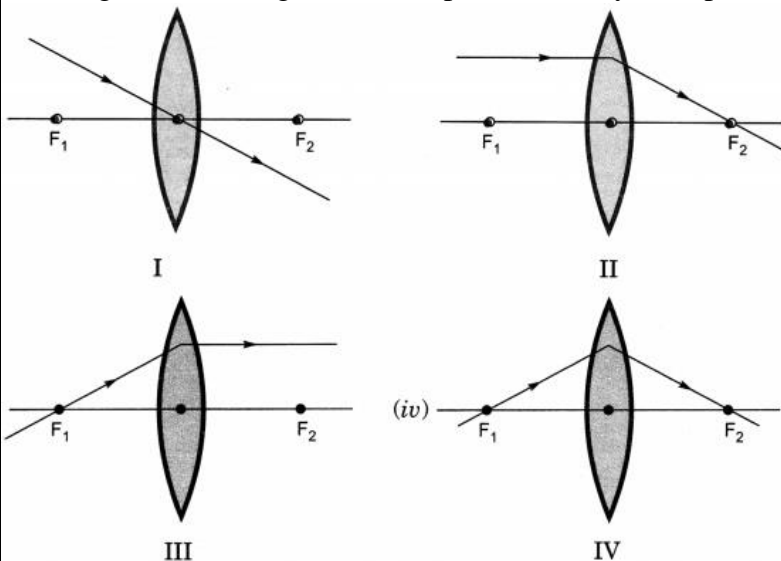


What is W, X, Y and Z?

|     | Bronchus | Bronchiole | Larynx | Trachea |
|-----|----------|------------|--------|---------|
| (a) | W        | X          | Z      | Y       |
| (b) | X        | Z          | Y      | W       |
| (c) | Y        | W          | X      | Z       |
| (d) | Z        | Y          | W      | X       |

42.

The diagrams showing the correct path of the ray after passing through the



- (a) II and III only  
 (b) I and II only  
 (c) I, II and III  
 (d) I, II and IV

43.

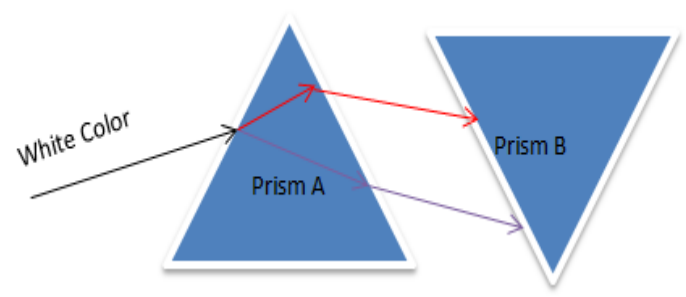
An object is placed in front of a convex lens of focal length 20 cm. The distance of the object from the lens is 30 cm. Find the position and magnification of the image.


- (a) 60 cm and -2  
 (b) -60 cm and -2  
 (c) 60 cm and +2  
 (d) 30 cm and -1

44.

Which of the following phenomena contributes significantly to the reddish appearance of the sun at sunrise or sunset?

- (a) Dispersion of light  
 (b) Scattering of light  
 (c) Total internal reflection of light  
 (d) Reflection of light from the earth

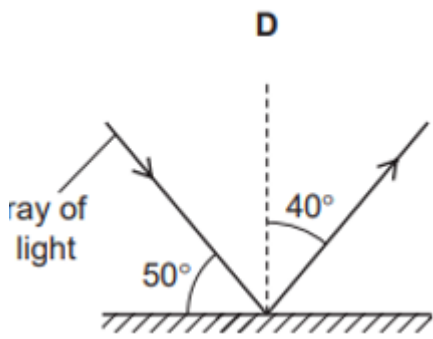
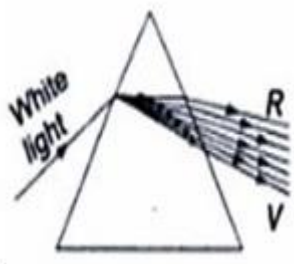
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| 45.   | <p>Consider the following situation, what can be said about emergent ray/rays?</p>  <p>(a) VIBGYOR<br/>(b) A Red light<br/>(c) A White Light<br/>(d) A violet light</p>   |
| 46.   | <p>Suppose you are in danger, and you need to give signal to save yourself, what colour you prefer?</p> <p>(a) Red<br/>(b) Yellow<br/>(c) Blue<br/>(d) Violet</p>   |
| 47.   | <p>Refractive index of water and benzene w.r.t air are 1.33 and 1.50 respectively. Calculate refractive index of benzene with respect to water?</p> <p>(a) 0.89      (b) 1.13      (c) 2.14      (d) 0.67</p>   |
| 48.   | <p>Which of the following phenomena of light are involved in the formation of a rainbow?</p> <p>(a) Reflection, scattering and dispersion<br/>(b) Refraction, dispersion and scattering<br/>(c) Refraction, dispersion and internal reflection<br/>(d) Dispersion, scattering and total internal reflection</p>   |
| <p><b>SECTION – C</b></p> <p>Section- C consists of three Cases followed by questions. There are a total of 12 questions in this section. Attempt any 10 questions from this section.<br/><u>The first attempted 10 questions would be evaluated.</u></p> |   |
| <b>CASE</b>   | <p>A scale for measuring hydrogen ion concentration in a solution, called pH scale has been developed. The p in pH stands for ‘potenz’ in German, meaning power. On the pH scale we can measure pH from 0 to 14. pH should be thought of simply as a number which indicates the acidic or basic nature of a solution. Higher the hydronium ion concentration, lower is the pH value</p> |

| 49.         | <p>Which one of the following will have the highest hydrogen ion concentration?<br/>         (a) pH = 1.1 (b) pH = 2.2 (c) pH = 3.3 (d) pH = 4.4</p>   |          |          |   |   |   |    |   |   |   |    |
|-------------|--|----------|----------|---|---|---|----|---|---|---|----|
| 50.         | <p>How is the hydrogen ion concentration and pH related to each other?<br/>         (a) They are inversely proportional.<br/>         (b) They are directly proportional.<br/>         (c) They are equal.<br/>         (d) They have no relation.</p>   |          |          |   |   |   |    |   |   |   |    |
| 51.         | <p>A basic solution could have a pH of:<br/>         (a) 3 (b) 5 (c) 7 (d) 9</p>   |          |          |   |   |   |    |   |   |   |    |
| 52.         | <p>The table provides the pH of four solutions A, B, C and D</p> <table border="1" data-bbox="683 926 1044 1117"> <thead> <tr> <th>Solution</th> <th>pH value</th> </tr> </thead> <tbody> <tr> <td>A</td> <td>1</td> </tr> <tr> <td>B</td> <td>10</td> </tr> <tr> <td>C</td> <td>4</td> </tr> <tr> <td>D</td> <td>12</td> </tr> </tbody> </table> <p>Which of the following correctly represents the solutions in increasing order of their hydronium ion concentration?<br/>         (a) <math>A &gt; B &gt; C &gt; D</math><br/>         (b) <math>A &gt; D &gt; B &gt; C</math><br/>         (c) <math>D &lt; B &lt; C &lt; A</math><br/>         (d) <math>D &lt; A &lt; B &lt; C</math></p> | Solution | pH value | A | 1 | B | 10 | C | 4 | D | 12 |
| Solution    | pH value   |          |          |   |   |   |    |   |   |   |    |
| A           | 1  |          |          |   |   |   |    |   |   |   |    |
| B           | 10   |          |          |   |   |   |    |   |   |   |    |
| C           | 4  |          |          |   |   |   |    |   |   |   |    |
| D           | 12   |          |          |   |   |   |    |   |   |   |    |
| <b>CASE</b> | <p><b>RESPIRATION</b></p>  <p><i>Graph showing Respiratory rate of different Activities</i></p> <p>Most living things need oxygen to obtain energy from food. The oxygen reacts with food</p>  |          |          |   |   |   |    |   |   |   |    |

|             |   |
|-------------|---|
|             | <p>molecules and that is how energy is obtained which is stored in the form of ATP molecules in the cells. This energy can be used anywhere the body wants to do so. The process of releasing energy from food, is called respiration.</p>  |
| 53.         | <p>What is the full form of ATP?<br/>         (a) Adenosine tri-phosphate<br/>         (b) Adenosine tri-phosphate<br/>         (c) Adenosine tetraphosphate<br/>         (d) Adenosine monophosphate</p>   |
| 54.         | <p>Respiration is<br/>         (a) Biochemical process<br/>         (b) Passive process<br/>         (c) Physical process<br/>         (d) Biophysical process</p>  |
| 55.         | <p>Respiration is the process in which-<br/>         (a) Energy is used up<br/>         (b) Energy is stored in the form of ADP<br/>         (c) Energy is not released at all<br/>         (d) Energy is released and stored in the form of ATP</p>  |
| 56.         | <p>The form of energy used in respiration is -<br/>         (1) Electrical energy<br/>         (2) Chemical energy<br/>         (3) Mechanical energy<br/>         (4) Radiant energy</p>   |
| <b>CASE</b> | <p>Ram placed a glass slab on a drawing sheet and marked its boundary as PQRS. Then he removed the glass slab and drew a line AO on the side PQ. He placed the glass slab in position, passed light from a laser torch through it along AO. He observed the path of light through the glass slab and marked the points O, B and C. PQ is the surface of separation of air and glass and that RS is the surface of separation of glass and air</p> |

|     |  |
|-----|--|
| 57. | Where does the incident and emergent ray meet in above figure?<br>a) inside the glass b) infinity c) above the glass slab d) none of these               |
| 58. | Which is of greater optical density?<br>a) Air<br>b) Glass<br>c) Both Glass and Air are optical denser<br>d) Both Glass and Air are optical rarer        |
| 59. | Is the angle of refraction greater or lower than the angle of incidence when it goes from glass to air?<br>a) Greater b) Lower c) Equal d) None of these |
| 60. | In the above figure lateral displacement is mentioned by<br>a) AO b) OB c) BC d) CD  |

| Q.NO | ANSWERS   |
|------|---|
|      | Section - A   |
| 1.   | (b) White to grey   |
| 2.   | (a) Hydrogen and copper oxide respectively                    |
| 3.   | (b) Treating the water with vinegar(acidic)                   |
| 4.   | (b) Basic oxide   |
| 5.   | (c) H <sub>2</sub> SO <sub>4</sub>                            |
| 6.   | c) combining with water                                       |
| 7.   | (c) colourless  |
| 8.   | (a) Ammonium chloride (acidic salt)                           |
| 9.   | (b) (ii)  |
| 10.  | (b) Brown (copper is deposited)                               |
| 11.  | . (a) Heterotrophs do not synthesis their own food            |
| 12.  | (a) Lime water turns milky due to the CO <sub>2</sub> exhaled |
| 13.  | (c) selective reabsorption                                    |
| 14.  | (a) Carbon dioxide  |
| 15.  | (c)(i), (ii) & (iii)  |
| 16.  | (d) (ii) and (iv)   |
| 17.  | c) C (angle of incidence =angle of emergence)                 |
| 18.  | (b) Bends towards normal                                      |

|                    |   |
|--------------------|---|
| 19.                | (b) MS (focal length- distance between focus and the pole)  |
| 20.                | D (angle of incidence-x and angle of refraction-z)  |
| 21.                | D (angle of incidence= angle of reflection)<br>  |
| 22.                | C (Object should be placed between 2F and F)  |
| 23.                | A.<br>  |
| 24.                | 1. b)-75cm<br>$m = -\frac{v}{u} \Rightarrow 1.5 = -\frac{v}{-25}$ $\Rightarrow v = \frac{75}{2} \text{ cm}$ <p>Now, from mirror formula,</p> $\frac{1}{f} = \frac{1}{v} + \frac{1}{u} = \frac{1}{75/2} + \frac{1}{-25} = -\frac{1}{75}$ $\therefore f = -75 \text{ cm}$ |
| <b>Section - B</b> |   |
| 25.                | (d) the solution remains colourless and no deposition is seen on the iron nail.(no displacement reaction as aluminium is more reactive than iron)   |
| 26.                | (c) displacement reaction   |
| 27.                | (c) both (a) and (b)  |
| 28.                | (b) brown residue   |



|     |  |
|-----|--|
| 29. | (b) Red litmus   |
| 30. | d) Weak acid and strong base   |
| 31. | A Both A and R are correct and R is the correct explanation of assertion.  |
| 32. | A Both A and R are true and R is the correct explanation of A  |
| 33. | E. Both A and R are true and R is the correct explanation of A   |
| 34. | 2. (c) Assertion is true and Reason is false   |
| 35. | (c) Cu   |
| 36. | (d) B and D  |
| 37. | c) (ii), (i), (iii), (iv)  |
| 38. | <b>d)</b> Glucose $\xrightarrow{\text{Cytoplasm}}$ Pyruvate + Energy $\xrightarrow{\text{Mitochondria}}$ CO <sub>2</sub> + H <sub>2</sub> O + Energy |
| 39. | b) Lime water  |

|     |   |
|-----|---|
| 40. | a) a) (ii), b) (i), c) (iv) d) (iii)  |
| 41. | (d) Z Y W X   |
| 42. | 3. (c) I, II and III  |
| 43. | (a) 60 cm and -2  |
| 44. | (b) Scattering of light   |
| 45. | (c) A White Light   |
| 46. | a) Red<br>—   |
| 47. | (b) 1.13<br><i>Refractive index of water with respect to air = <math>\mu_w = 1.33</math></i><br><i>Refractive index of benzene with respect to air = <math>\mu_B = 1.50</math></i><br><i>Refractive index of benzene with respect to water = <math>\frac{\mu_B}{\mu_w} = \frac{1.5}{1.33} = 1.13</math></i> |

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| 48.                | 4. (c) Refraction, dispersion and internal reflection             |
| <b>Section - C</b> |   |
| 49.                | (a) pH = 1.1(lower the pH, higher the hydrogen ion concentration) |
| 50.                | (a) They are inversely proportional.                              |
| 51.                | (d) 9(basic solutions have pH more than 7)                        |
| 52.                | (c) $D < B < C < A$   |
| 53.                | (a) Adenosine tri phosphate                                       |
| 54.                | a) Biochemical process  |
| 55.                | (d) Energy is released and stored in the form of ATP              |
| 56.                | (2) Chemical energy   |
| 57.                | 5. b) infinity  |
| 58.                | 6. b) Glass<br>_____  |
| 59.                | 7. a) Greater<br>_____  |
| 60.                | 8. d) CD  |

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